



$2\text{H}_2(g) + \text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(g)$ combination (synthesis)
combustion
oxidation-reduction

$\text{HNO}_3(aq) + \text{KOH}(aq) \rightarrow \text{KNO}_3(aq) + \text{H}_2\text{O}(l)$ acid-base
double-replacement

$\text{Pb}(\text{NO}_3)_2(aq) + \text{KI}(aq) \rightarrow \text{PbI}_2(s) + \text{KNO}_3(aq)$ precipitation
double-replacement

$\text{KClO}_3(s) \rightarrow \text{KCl}(s) + \text{O}_2(g)$ decomposition
oxidation-reduction

$\text{C}_6\text{H}_{12}\text{O}_6(s) + 6\text{O}_2(g) \rightarrow 6\text{CO}_2(g) + 6\text{H}_2\text{O}(g)$ oxidation-reduction
combustion

$2\text{H}_2\text{O}(g) \rightarrow 2\text{H}_2(g) + \text{O}_2(g)$ decomposition
oxidation-reduction

$\text{Zn}(s) + 2\text{HCl}(aq) \rightarrow \text{ZnCl}_2(aq) + \text{H}_2(g)$ single-replacement
oxidation-reduction

$2\text{Zn}(s) + \text{O}_2(g) \rightarrow 2\text{ZnO}(s)$ oxidation-reduction
combination (synthesis)
combustion