

# Acids & Bases in Aqueous Media (25°C)

$\text{pH} = -\log[\text{H}^+]$	$[\text{H}^+] = 10^{-\text{pH}}$
$\text{pOH} = -\log[\text{OH}^-]$	$[\text{OH}^-] = 10^{-\text{pOH}}$
$\text{pH} + \text{pOH} = 14$	$[\text{H}^+][\text{OH}^-] = 10^{-14}$

