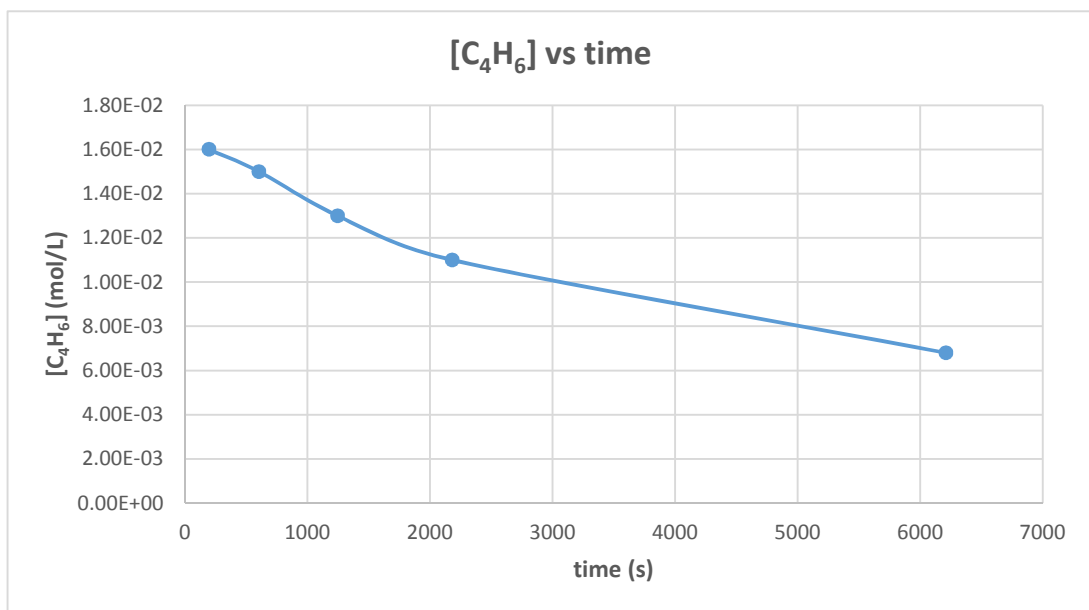


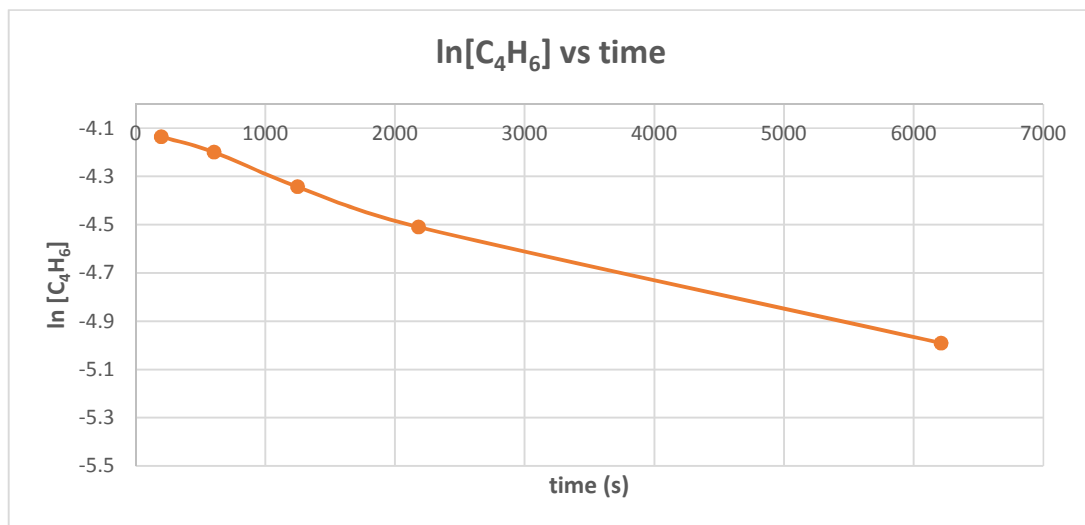
Time (s)	[C <sub>4</sub> H <sub>6</sub> ] (mol/L)	ln[C <sub>4</sub> H <sub>6</sub> ]	1/[C <sub>4</sub> H <sub>6</sub> ] (L/mol)
195	1.60E-02	-4.13517	62.50
604	1.50E-02	-4.19971	66.67
1246	1.30E-02	-4.34281	76.92
2180	1.10E-02	-4.50986	90.91
6210	6.80E-03	-4.99083	147.06

Scroll down  
to see  
all graphs

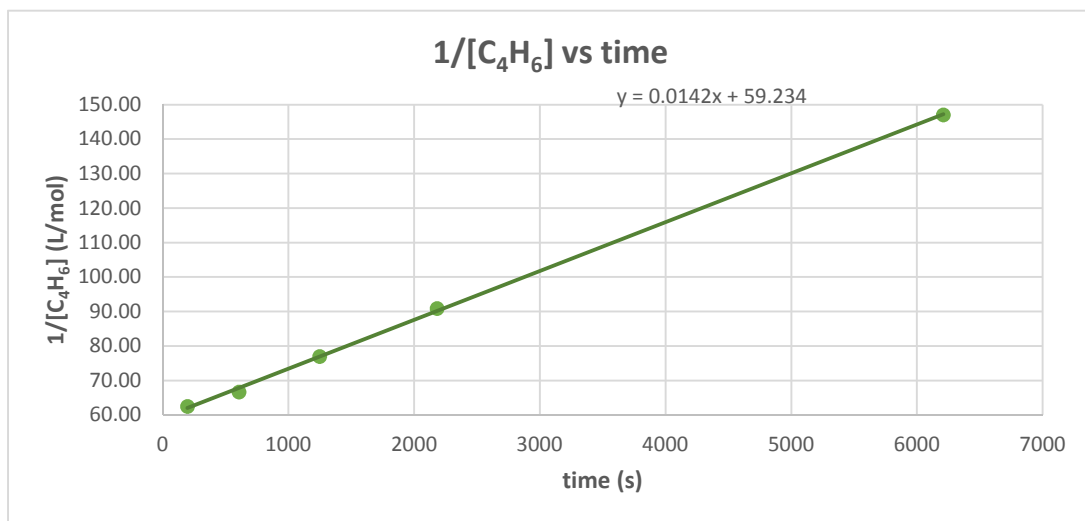




Since the plot [C<sub>4</sub>H<sub>6</sub>] vs time is *not* a line, the reaction is *not* zero order.



Since the plot  $\ln[C_4H_6]$  vs time is *not* a line, the reaction is *not* first order.



Since the plot  
**1/[C<sub>4</sub>H<sub>6</sub>] vs time IS a  
line, the reaction IS  
zero order.**