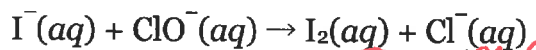


CHEM 101B - Balancing Redox Reactions in Acidic or Basic Solution

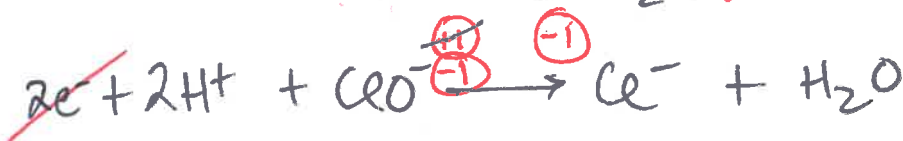
1. Balance the following oxidation-reduction reaction that occurs in acidic solution:



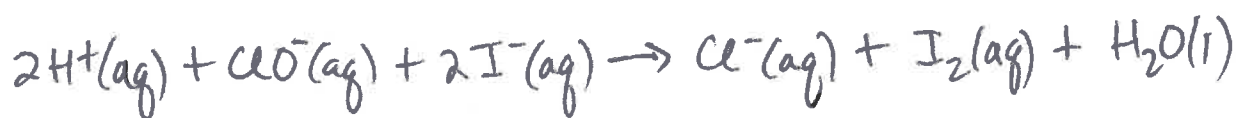
OX



red



$$1 - (-1) = 2$$



check

$$\text{H} \quad 2, 2 \quad \checkmark$$

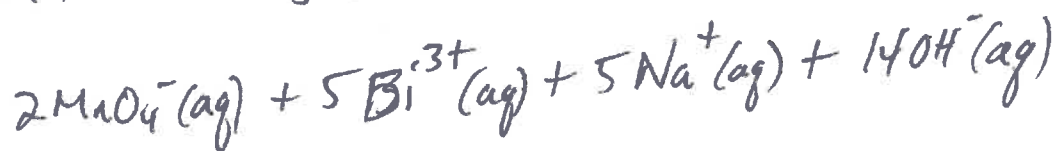
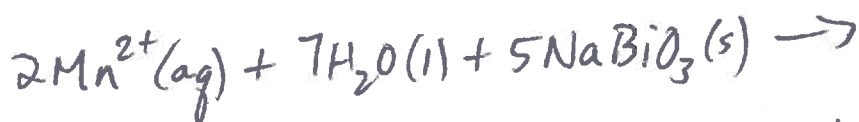
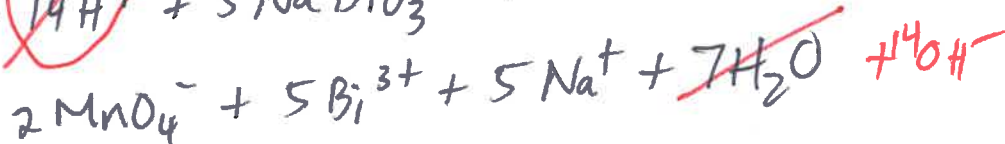
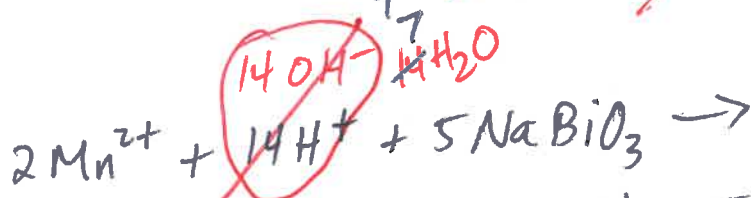
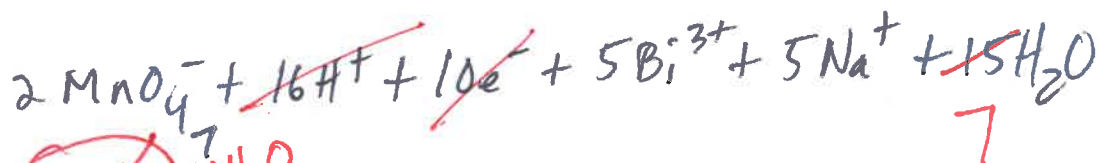
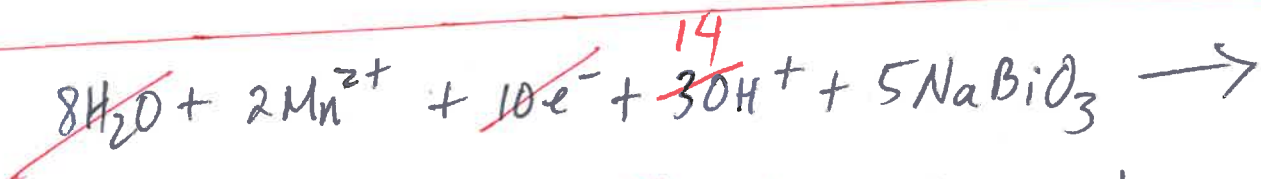
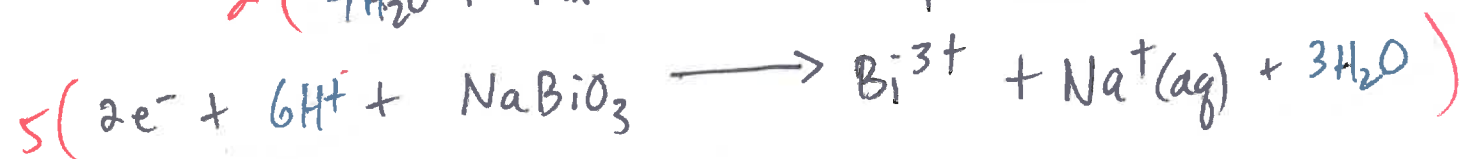
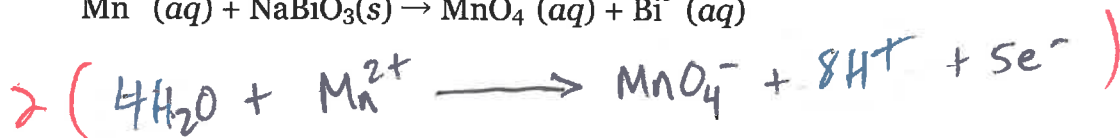
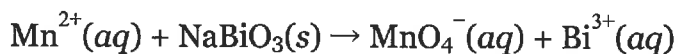
$$\text{Cl} \quad 1, 1 \quad \checkmark$$

$$\text{O} \quad 1, 1 \quad \checkmark$$

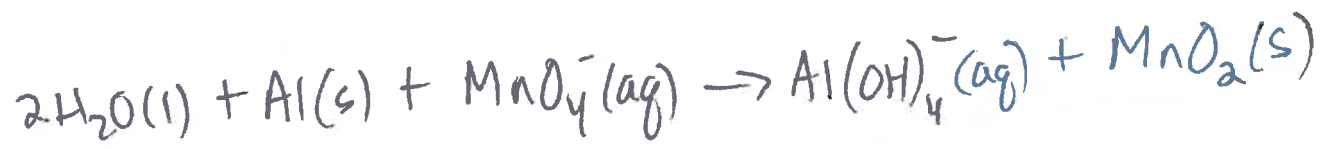
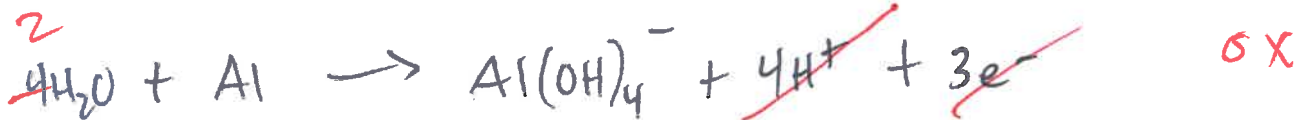
$$\text{I} \quad 2, 2 \quad \checkmark$$

$$\text{charge} \quad -1, -1 \quad \checkmark$$

2. Balance the following oxidation-reduction reaction that occurs in basic solution:



3. Balance the following oxidation-reduction reaction that occurs in basic solution:



4. Balance the following oxidation-reduction reaction that occurs in acidic solution:

